Structural Characterization of a Cationic Zirconocene Olefin Polymerization Catalyst with its Methylated Boralumoxane Counterion**

Bodo Richter, Auke Meetsma, Bart Hessen,* and Jan H. Teuben

The discovery by Sinn and Kaminsky that methyl alumoxane (MAO) can act as an efficient activator for metalloocene olefin polymerization catalysts[1] has triggered tremendous developments in single-site olefin polymerization catalysis.[2] Presently, more than twenty years after the initial discovery, the actual nature of MAO and its mechanism of catalyst activation are still under debate.[3] Model studies on tert-butyl alumoxanes led Barron et al. to formulate the hypothesis that these alumoxanes consist of oligomeric (RAIO)n clusters, containing 4-coordinate Al in strained fused 4-membered Al2O2 rings that can exhibit “latent Lewis acidity” by ring opening, allowing Al to abstract an alkyl anion from the transition metal dialkyl catalyst precursor.[4, 5] The direct (structural, spectroscopic) study of these processes and their products is hampered by the apparent equilibrium nature of the alkyl transfer reaction, with the equilibrium constant strongly favoring the starting materials when well-defined alumoxanes with sterically demanding alkyl groups are used. Crystal structure determinations of the products resulting from the reaction of [tBuAlO], with MeLi[4] and with RNH,[6] gave support for the proposed activation mechanism, but as yet the products of the activation of metallocone single-site olefin polymerization catalysis by alumoxane activators have eluded full characterization.

Recently we reported the synthesis and structural characterization of a well-defined boralumoxane species, [tBu-Al,Ar,BiO] (1, Ar = 2,6-diisopropylphenyl), and showed that this compound is able to activate [Cp2ZrMe2] for catalytic ethene polymerization.[7] Although 1 is topologically quite different from the tBu-alumoxane species isolated by Barron et al., containing 3-coordinate B and 4- and 5-coordinate Al, it shares with these compounds the presence of a strained 4-membered Al2O2-ring assembly (edge sharing with a BAIO2

[1] Dr. B. Hessen, Dr. B. Richter, A. Meetsma, Prof. Dr. J. H. Teuben Dutch Polymer Institute/Center for Catalytic Olefin Polymerization Stratingh Institute for Chemistry and Chemical Engineering University of Groningen Nijenborgh 4, 9747 AG Groningen (The Netherlands) Fax: (+31) 50-3634315 E-mail: hessen@chem.rug.nl

[2] Netherlands Institute for Catalysis Research (NIOK) publication RUG 02-04-01.

[3] Supporting information for this article is available on the WWW under http://www.angewandte.org or from the author.
Lewis acidity. The 1H NMR spectrum of Angew. Chem. Int. Ed. 2002 2 shows this to be the salt [Cp*2 ZrMe2] with pyridine to give 2.

Figure 1. Molecular structure of 2. All carbon and hydrogen atoms omitted, except for the ipso carbons of the groups directly bound to Al (iBu) or B (2,6-Pr) in 59% yield. A crystal structure determination[9] of 2-(C6H5)3 obtained by crystallization from benzene, showed a cluster (Figure 1) with basic structural features similar to those of 1, including a central 5-coordinate Al atom and interconnected 6-membered and 4-membered Al/B-O rings, but without the strained Al2O2/BAlO2 edge-sharing ring assembly. The pyridine molecule is coordinated to one of the two iBuAl moieties. It is possible that this structure arises from initial attack of pyridine on the iBuAl group in the strained ring, followed by a rearrangement of the oxygen atoms around the 5-coordinate Al core, as shown in Scheme 1. These intramolecular rearrangements are likely to be facile in the boralumoxane framework, as was shown by the rapid fluxionality of 1, resulting in solution NMR spectra consistent with an average C2 symmetry for this compound.[17] The formation of 2 from 1 and pyridine thus appears to be in line with the expected reactivity of 1 based on the concept of latent Lewis acidity. The 1H NMR spectrum of 2 is complex, but clearly shows the resonances of the coordinated pyridine, as well as four separate resonances for the AlMeBu groups, indicating an asymmetric structure.

Adding one equivalent of [Cp*2 ZrMe2] (Cp* = η5-C5Me5) to a solution of 1 in C6D6 in an NMR tube at ambient temperature results in a 1H NMR spectrum that is essentially a superposition of the spectra of the starting materials. This suggests that with this zirconocene the reaction of the boralumoxane is either very sluggish, or invisible due to a highly unfavorable equilibrium constant for a possible reversible methyl transfer. In order to trap the [Cp*2 ZrMe2] cation that would result from such a methyl transfer reaction, we added an excess of allyl methyl thioether to the solution containing 1 and [Cp*2 ZrMe2]. Allyl methyl thioether does not react with 1 itself, but is known to give stoichiometric insertion of the olefin into the Zr–Me bond of [Cp*2 ZrMe2], resulting in the cationic chelate complex [Cp*2 ZrCH2CH=MeSMe]– (its OEt analogue was structurally characterized previously in our group[9]). The result of this reaction was a gradual precipitation over several hours of a red oily material and the disappearance of the NMR resonances associated with 1 and [Cp*2 ZrMe2].

Permuting the reaction on a preparative scale in a toluene/pentane mixture resulted in the formation of red-orange crystalline material. A crystal structure determination[9] showed this to be the salt [Cp*2 ZrCH2CH(Me)CH2SMe]-, [MerBu4Al4Ar4B4O8][Me]3, cocrystallized with 1.5 equivalents of toluene per formula unit. The material was obtained in 89% yield. In Figure 2 top, cation and anion are shown schematically, revealing the position of the methyl group that has been transferred from the zirconocene, and in Figure 2 bottom, the core structure of the anion in 3 is shown. As with the pyridine reaction product 2, the newly introduced group is attached to one of the BuAl moieties, and again the strained Al2O2/BAlO2 edge-sharing ring assembly has been opened. The main difference with the structure of 2 is that the rearrangement after ring opening leads to a different arrangement around the central Al atom, but again this structure is readily derived from the presumed initial product of the ring opening (Scheme 2).

Performing the reaction in C6D5Br and following it with 1H NMR spectroscopy showed the gradual formation of the known cation [Cp*2 ZrCH2CH(Me)CH2SMe]+, and in addition a complicated set of resonances for the methylated boralumoxane anion appeared. Initially, a resonance appeared at δ = −0.58 ppm, which is likely to be associated with the methyl group that is transferred to Al. In the course of time (hours) this species transformed into a product with the AlMe resonance at δ = −0.40 ppm, whereas the resonances of...
the cation remained unchanged. This suggests that the methylated boralumoxane anion is converted from a kinetically controlled to a thermodynamically controlled product. For both isomers, four separate resonances for the AlMe groups are observed, indicating asymmetric structures. Dissolving a portion of crystalline 3 (obtained from the reaction in toluene/pentane as described above) in C₆D₆Br initially showed a dominant AlMe ¹H NMR resonance at δ = −0.58 ppm and subsequently a gradual dominance of the species with the resonance at δ = −0.40 ppm. This suggests that the crystal structure of 3 most likely represents the structure of the kinetically controlled product.

The present study conclusively shows that the strained AlₓOₓ/BAlₓOₓ edge-sharing ring assembly in the boralumoxane 1 is the feature responsible for the reactivity of this compound towards Lewis bases as well as zirconocene dimethyl complexes. Zirconocene dialkyl compounds are activated by alkyl transfer to the boralumoxane, and the resulting cationic species can insert olefins into the Zr—bond, as was seen by characterization of the product of the methyl allyl thioether insertion. These results give strong support for the concept of latent Lewis acidity formulated by Barron et al. for the mechanism of activation of zirconocene-based olefin polymerization catalysts by alumoxane activators. The observation of this reactivity with boralumoxane 1, which is topologically quite different from the [B₅Al₄O₅], alumoxanes, suggests that there may be a wider scope for metalacyclic activators for single-site olefin polymerization catalysts. It may be noted that the reaction between transition metal dialkyl compounds and alumoxane species stabilized with sterically demanding groups can be rather slow (for [Cp₂-ZrMe₂] and 1 in the order of hours at ambient temperature), making it more difficult to estimate the true activating potential of these species from polymerization experiments started by simple mixing of the metal dialkyl compound and the activator.

**Experimental Section**

2: Pyridine (26.5 μL, 0.328 mmol) was added to a suspension of 1 (0.378 g, 0.328 mmol) in 5 mL of n-pentane. Subsequently toluene (3 mL) was added, and the mixture was briefly warmed to 40 °C. After allowing the mixture to cool to ambient temperature, the supernatant was decanted and the precipitate was rinsed with 5 mL of toluene. The mixture was cooled to ambient temperature, the supernatant was decanted and the precipitate was rinsed with 5 mL of n-hexane. The precipitate was rinsed with 5 mL of n-hexane. Drying in vacuo yielded 2 (0.240 g, 59%). Selected NMR data: ¹H NMR (500 MHz, [D₆]benzene, 25 °C): δ = 7.52 (d, J = 5.0 Hz, 2H, py-o-H), 6.79 (t, J = 7.7 Hz, 1H, py-p-H), 6.40 (m, 2H, py-m-H), 1.26, 0.96, 0.83, 0.57 (s, 9H each, tBu) ppm. ¹³B NMR (160 MHz, [D₅]benzene, 25 °C): δ = 30.0 (W₂₁ = 2640 Hz) ppm. ²⁷Al (130 MHz, [D₅]benzene, 25 °C): δ = 120 (very br, W₂₁ > 50 kHz), 41.9 (W₂₁ = 5900 Hz) ppm. C,H,N analysis (%): calcd for C₉₀H₁₄₈O₈SB₄Al₄Zr(C₇H₈)₁₅ (1770.8): C 68.49, H 9.02; found C 68.38, H 8.88, N 1.07.

3: A solution of [Cp₂-ZrMe₂] (0.134 g, 0.343 mmol) in 2 mL of toluene was added to a suspension of 1 (0.368 g, 0.319 mmol) in 3 mL of n-pentane. Subsequently allyl methyl thioether (0.2 mL) was added, and the mixture was allowed to stand overnight at ambient temperature. The supernatant was decanted from the red-orange crystalline material, which was rinsed with two portions of n-hexane (3 mL each). Drying in vacuo yielded 3·(C₆H₆)₁₂ (0.505 g, 0.285 mmol, 89%). Selected NMR data: ¹H NMR (500 MHz, [D₆]bromobenzene, 25 °C): cationic part: δ = 2.3 (br, nH, C,HMe), 2.26 (ddd, J = 10.8, 4.3, 3.3 Hz, 1H, SCH₂), 2.14 (dd, J = 10.7, 12.6 Hz, 1H, CH₂), 2.14 (dd, J = 12.9, 12.3, 1H, CH₂), 1.94 (J = 12.9, 12.3, 1H, CH₂), 1.67 (s, 3H, CH₃), 1.60, 1.57 (s, 3H each, C,HMe), 0.79 (d, J = 6.3 Hz, 3H, CH₃), −1.12 (ddd, J = 12.9, 6.7, 3.4 Hz, 1H, ZrCH₂) ppm. anionic part, thermodynamic product: δ = 1.25, 1.14, 1.09, 0.41 (s, 9H each, tBu), −0.40 (s, 3H, AlMe) ppm. ¹³B NMR (160 MHz, [D₅]chloroform, 25 °C): δ = 28.6 (W₂₁ = 1900 Hz) ppm. ²⁷Al NMR (130 MHz, [D₅]chloroform, 25 °C): δ = 124 (very br, W₂₁ = 35 kHz), 42.5 (W₂₁ = 3600 Hz) ppm. C,H,N analysis (%): calcd for C₉₀H₁₄₈O₈SB₄Al₄Zr(C₇H₈)₁₂ (1770.8): C 68.49, H 9.02; found C 68.31, H 8.97.


[8] Crystallographic structure determinations were performed on a Bruker SMART Apex CCD diffractometer using MoKα radiation (λ = 0.71073 Å). The structures were solved with Patterson methods and extended by direct methods applied to difference structure factors. Crystallographic data for 2: CaH3AlBNo3 (C4H2). M1 = 1309.91, monoclinic, space group P21/n, a = 13.8868(6), b = 25.188(1), c = 23.346(1) Å, β = 106.805(1)°, V = 7817.2(6) Å3, Z = 4, P121 = 1.113 g cm−3, μ = 1.1 cm−1, T = 110 K. A final refinement on F2 converged at wR(F2) = 0.1789 for 13782 reflections and R(F) = 0.0733 for 7505 reflections with F2 ≥ 4σ(F2) and 857 parameters. Crystallographic data for 3: [CaH2S2][CaH6AlB4O12] (CaH2). M1 = 1770.83, monoclinic, space group P21/c, a = 20.629(1), b = 20.3125(1), c = 26.301(1) Å, β = 110.155(1)°, V = 10346.0(7) Å3, Z = 4, P121 = 1.137 g cm−3, μ = 2.11 cm−1, T = 110 K. A final refinement on F2 converged at wR(F2) = 0.0289 for 18278 reflections and R(F) = 0.0808 for 11143 reflections with F2 ≥ 4σ(F2) and 1145 parameters and 16 restraints. One of the toluene molecules in the lattice shows extensive orientational disorder, resulting in unrealistic displacement parameters when allowed to vary anisotropically. The other toluene molecule is ordered over an inversion center, and was described as a single molecule with 0.5 occupancy with geometrical restraints. CCDC-177577 (2) and CCDC-177576 (3) contain the supplementary crystallographic data for this paper. These data can be obtained free of charge via www.ccdc.cam.ac.uk/conts/retrieving.html (or from the Cambridge Crystallographic Data Centre, 12, Union Road, Cambridge CB21EZ, UK; fax: (+44) 1223-336-333; or deposit@ccdc.cam.ac.uk).


**BuMgNiPr2: A New Base for Stoichiometric, Position-Selective Deprotonation of Cyclopropane Carboxamides and Other Weak CH Acids**

Mao-Xi Zhang and Philip E. Eaton*

Organolithium (RLi) and -magnesium (R2Mg) compounds are kinetically poor bases for proton removal from weak carbon acids. If an amine like N,N,N,N-tetramethylethylendiamine (TMEDA) is added, the barrier is sometimes lowered, but the nucleophilicity of the organometallic compound remains a problem. This can be ameliorated by using metal amides like lithium disopropylamide (LDA) and bis(disopropylamido)magnesium (DA2Mg). These poor nucleophiles are still kinetically effective bases for deprotonation of weakly acidic CH groups (pKα ≈ 30–35).[1] However, as the pKα of the liberated amine and that of the CH acid (e.g. I) are similar, such deprotonations are nowhere near stoichiometric (e.g. Scheme 1[2]). This is unsatisfactory.

![Scheme 1](image)

We now introduce alkylmagnesium amides, here specifically BuMgNiPr2, denoted hereafter as BuMgDA, as an effective solution to this problem.[3] We prepare BuMgDA simply by adding 1.0 equivalent of anhydrous disopropylamine (DAH) to commercial dibutylmagnesium in heptane (ca. 1.0 M) at room temperature and then stirring the solution for five minutes at 50 °C. BuMgDA[3b] in heptane is quite reactive. Replacing heptane, all or in part, with THF (after the base has been formed) increases this usefully. Unlike BuMg and many organolithium bases, BuMgDA is stable even in refluxing THF for many hours. BuMgDA, like DA2Mg, deprotonates/metallates amide-activated cyclopropane-CH (a, β, or beyond) and cetyl-CH (ortho),[8] but the BuMgDA deprotonations are driven to completion by irreversible formation of butane.

It is instructive to compare the metalation of the cyclopropylcarboxamide 2 using BuMg solutions in heptane treated first with 0, 0.5, or 1.0 equivalent of DAH. BuMg itself reacted only slowly,[7] mostly starting material was recovered. In the other two cases, when at least some BuMgDA[3b] was present, the overall deprotonation/metalation/carboxylation proceeded in high yield, but the final

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**COMMUNICATIONS**


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